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(54) **Process for the conversion of olefins having from 4 to 12 carbon atoms to propylene.**

(57) A process for the conversion of olefins having from 4 to 12 carbon atoms into propylene, characterized in that said olefins are brought into contact with a zeolitic compound (optionally in admixture with a binder), selected from silicalites, borasilites, chromosilicates and those zeolites ZSM5 and ZSM11 in which $\text{SiO}_2/\text{Al}_2\text{O}_3$ (by moles) is > 350 , at a space velocity of from 8 to 200 kg/h of olefins per kg of pure zeolitic compound (binder excluded) and at a temperature of from 400 to 600°C, said zeolitic compound being used as such or in a modified form.

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PROCESS FOR THE CONVERSION OF OLEFINS HAVING FROM 4 TO 12
CARBON ATOMS TO PROPYLENE

The invention concerns a process for the conversion
5 of olefins having from 4 to 12 carbon atoms into propylene.

Nowadays huge amounts of olefine cuts, from C_4 to
 C_{12} , linear or branched, are available throughout the
world and they are widely employed for different purposes,
such as described for instance in Italian patent publi-
10 cations 24152 A/82, 24550 A/82 and 19292 A/83. Sometimes,
however, because of contingent reasons, even outside the
chemical field, like for instance transportation difficul-
ties, it would be better to have still further possibilities
of use. A promising use of said olefins would be their
15 conversion into propylene and/or ethylene.

However, endurance tests have shown that the excellent
initial behavior of some catalysts, like for instance ZSM5
and ZSM11, disappears after some time; after a few weeks
conversion and selectivity drop to poor levels.

20 An object of the invention is to provide a process for
catalytically converting olefin cuts C_4 - C_{12} with high
propylene (and optionally also ethylene) yields and with
a long life time of the catalyst before substitution or
regeneration.

25 In its most general form, the invention concerns a
process for the conversion of olefins having from 4 to 12
carbon atoms into propylene which is characterized in that
said olefins are brought into contact with a zeolitic com-
pound (optionally in admixture with a binder) selected
30 from silicalites, boralites, chromosilicates and those
zeolites ZSM5 and ZSM11 in which (SiO_2/Al_2O_3) (by moles) is
 ≥ 350 , at a space velocity of from 5 to 200 kg/h of olefins
per kg of pure zeolitic compound, binder excluded, and at
a temperature of from 400 to 600°C, said zeolitic compound
35 being used as such or in a modified form.

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The behavior of silicalites depends on the conversion pressure; for instance, if the silicalite is silicalite-1 and if the pressure is substantially atmospheric, the space velocity must be lower than 50 h^{-1} ; if said compound is silicalite-1 and if the pressure is from 1.5 to 7.5 absolute atmospheres the space velocity must be in general above 50 h^{-1} .

Best results are obtained when said catalytic silicalite-1 is activated in the conversion reaction of the olefins into propylene, under said operative conditions. This initial (activating) run takes at least 8 and preferably 12 hours; the silicalite-1 is used as such or in a modified form and the modifying element is selected from Cr, Mg, Ca, Sr and Ba.

The modifying element can be incorporated into the catalyst by means of ion exchange or by any other method, for instance impregnation or co-precipitation during the synthesis of the zeolitic compound.

According to a particularly advantageous way for the preparation of a non-modified silicalite, the raw product coming from the zeolite synthesis is dried, for instance at 120°C , calcined (e.g. at 540°C for some hours), in order to remove all the residual organic templating agent, and then exchanged with an aqueous solution of HCl , NH_4Cl , NH_4NO_3 or an equivalent H^+ or NH_4^+ source. When an ammonium compound is used, it is necessary to heat, e.g. at 400°C , in order to obtain the acid form of the silicate. A survey of techniques alternative to ion exchange is given e.g. in U.S. patents 3,140,249; 3,140,251; 3,140,253 and in European patent publications 30796, 36707, 37168, 40463, 68754.

The zeolitic material, after calcination and conversion into the acid form, shows a long endurance and a very high catalytic activity. These zeolitic compounds can be used as catalysts (as such or in a modified form) optionally in admixture with suitable amounts of binders, for instance SiO_2 or Al_2O_3 . A list of other binders can be found

e.g. in European patent publication 36707. The regeneration can be carried out in air for some hours, at 400-600°C. A
 5 steam regeneration is described in European patent publication 36704 and according to a further and very successful method, the catalyst can be regenerated by a hydrogen treatment.

As to the initial activation of the catalyst, some
 10 methods are described in European patent publications 34444 and 35830; in general it is advisable to activate the catalyst for some hours in air, at 450-750°C (preferably 540-700°C). Furthermore the conversion itself (of olefins to propylene) has an activating effect on the zeolitic
 15 catalyst. In other words, catalyst and reaction affect each other in a mutual, reciprocal and beneficial activating reaction.

Any process for the conversion of more or less heavy olefinic cuts into propylene will be indicated hereafter,
 20 as a "post-pyrolysis" process. When the feed of a post-pyrolysis process is a mixture of olefins having 4 C atoms, there is a considerable problem to be solved, because the C₄ cuts always contain substantial amounts of paraffins, in general, also having 4 C atoms, which paraffins pass
 25 (at 400-600°C) the zeolitic bed without taking part in any reaction. Furthermore a small amount of C₄⁻ paraffins is produced by the post-pyrolysis process itself. The conversion to C₃H₆ could be enhanced by a recycle of the non-reacted C₄⁻olefins or of the C₄⁻olefins formed during the
 30 reaction. In such a case, however, an increasing accumulation of n-butane and of isobutane would take place. This drawback could be avoided by a separation of paraffins from olefins before feeding the reactor but such a separation is rather difficult. Butenes and isobutane cannot be isolated by a simple distillation and it is usually
 35 necessary to carry out an extractive distillation (a complicated technique), which is particularly burdensome for the C₄⁻ cuts coming from catalytic cracking, where butane and isobutane may account for even 50% of the whole. The problem, however, can be solved in a surprisingly easy way

by the process according to the invention, when employing an integrated oligomerization step. In other words, a particular embodiment of the invention (the feed being a paraffin-olefin mixture) comprises the following steps (reference is made also to figure 2):

- a) preliminary oligomerization of a ($C_4^- + C_4^+$) mixture at a temperature of from 320 to 380°C using a catalyst bed of zeolitic nature (see e.g. U.S. patent 4,150,062) to obtain a mixture of olefins having from 5 to 8 C atoms, with the C_4^+ paraffins remaining unconverted;
- b) cooling and condensation of the oligomerization effluent in order to separate the C_4^+ paraffins as a gaseous phase, and conversion of the remaining ($C_5^- - C_8^-$) mixture to propylene under typical "post-pyrolysis" operative conditions;
- c) cooling of the effluent from the reactor for the conversion to propylene and compression of said effluent, preferably at 13-16 absolute Kg/cm², whereby the hydrocarbons having 4 or more C atoms are condensed and the hydrocarbons having less than 4 C atoms are separated as a gaseous phase.

These hydrocarbons ($< C_4$) can be advantageously recycled to a conventional battery of distillation columns for thermal or catalytic cracking, in order to recover all the propylene contained therein. The small and possible amounts of aromatics (BTX) can be easily separated from the other C_4 hydrocarbons and recycled together with the final ($C_4^- - C_8^-$) mixture, containing small amounts of butanes produced during the conversion reaction described under item (b) above.

In order to carry out the oligomerization, C_4 olefins, containing C_4 paraffins in any proportion, are initially brought into contact with a catalyst of zeolitic nature, for instance ZSM5 or ZSM11, in an acid or in a modified form, at 250-400°C (preferably 320-380°C) and at space

5 velocities from 2 to 10 (preferably 4 - 8) kg/h of reactants per kg of pure zeolite (binder excluded). In other words, the olefins of the ($C_4^- + C_4^+$) mixture are converted almost totally into a ($C_5^- - C_8^-$) olefinic mixture, while butane and isobutane do not react. The separation of the butanes can thus be carried out very easily by simple cooling with water at room temperature. The olefinic ($C_5^- - C_8^-$) mixture liquifies while the butanes are separated as a gaseous phase, said ($C_5^- - C_8^-$) mixtures being optimal raw-materials for the production of propylene.

15 Depending on the operative conditions of the synthesis of the silicalite, as e.g. the dilution of the starting solutions, the resulting crystallites may have a widely variable size. The crystallites to be employed in the process according to the invention should have, in general, rather small sizes, such small sizes being obtainable, for instance, following the teachings of U.S. patent 3,926,782. 20 The following examples are given merely for illustration purposes and do not in any way limit the scope of the invention.

OPERATIVE CONDITIONS COMMON TO ALL EXAMPLES

As to the method of preparation of the catalysts see:

- 25 - for silicalite-1: U.S. patent 4,061,724;
- for boralite: Taramasso et al.: "Molecular Sieve Borosilicates", Proc. 5th Int. Conf. on Zeolites, Naples 1980, pages 40-48 (Heyden and Son Ltd. London 1980); the boralite used in the examples has a ratio $SiO_2 : B_2O_3$ of 7 (by moles); a more recent method for obtaining boro-silicates is described in European patent publication 77946 .
- 30 - for chromosilicates: Italian patent publication 22568 A/82, in the name of the Applicant; the chromosilicate used in the examples has a ratio $SiO_2 : Cr_2O_3$ of 38 (by moles).

35 As to zeolites ZSM5 showing a very high $SiO_2 : Al_2O_3$ ratio, not exemplified, see Italian patent publication 21699 A/83, in the name of the Applicant. In the absence of different indications, all the catalysts were activated 2 hours at 540°C before being used.

EXAMPLES 1-6

3 g of zeolitic catalyst, in admixture with 0.9 g of SiO_2 (as a binder), were loaded into a microreactor which was continuously fed, at a pressure slightly above atmospheric pressure with a 50/50 mixture of butene-2-trans and butene-2-cis. Operative conditions and results are given in Table 1.

EXAMPLE 7

Example 4 was repeated using a silicalite-1 impregnated with a chromium salt, thus obtaining slightly better results.

EXAMPLES 8-15

0.5 g of zeolitic catalyst, in admixture with 0.21 g of Al_2O_3 (as a binder), were loaded into a microreactor which was continuously fed, at a pressure slightly above the atmospheric pressure with a 70/100 mixture of n-pentene and of helium (by moles). Detailed operative conditions and results are given in Table 2. When modified zeolitic catalysts were used, the modifying element was added by means of impregnation.

EXAMPLES 16-23

Examples 8 and 9 were repeated replacing pentene with n-hexene (ex. 16 and 17), with 2-methyl-pentene-1 (ex. 18, 19 and 20) and with n-octene-1 (ex. 21, 22 and 23), respectively, the parameters being slightly modified as indicated in Table 3, which shows also the very good results.

EXAMPLES 24-29

2.12 g of zeolitic catalyst were loaded into a microreactor which was continuously fed with 7.5 kg/h of isobutene per kg of zeolitic catalyst at 550°C and at a pressure slightly above atmospheric pressure. Data and results are given in Table 4.

EXAMPLES 30 AND 31

Examples 24 and 25 were repeated raising the pressure to 5 absolute atmospheres and the space velocity (WHSV) from 7.5 to 45 h^{-1} , thus obtaining analogous results.

EXAMPLES 32-35

3 g of silicalite-1, without binder, were loaded into a microreactor which was continuously fed with 60kg/h of the olefins of example 1 per kg of silicalite, at high pressures (6 absolute atmospheres for example 32 and 33 and 8 abs. atm. for ex. 34 and 35). The detailed operative conditions and the results are given in Table 5. The results show clearly that a slight increase in pressure, corresponding approximately to the pressure of the olefins in the industrial manufacturing plants, allows the same conversions and selectivities, but at a much higher velocity. In other words, it is surprising and was quite unexpected that a suitable increase of the pressure increases very much the productivity of the silicate and therefore the output of the industrial plants.

EXAMPLE 36

~~50 parts by weight of silicalite-1 were admixed with~~ 50 parts of Al_2O_3 (binder) and the mixture was loaded into a microreactor which was continuously fed at atmospheric pressure, with a 50/50 mixture of butene-2-trans and butene-2-cis, at 550°C and at a space velocity of 20 kg/h of olefins per kg of silicalite (binder excluded). The run was very long (120 h) and the results, hour by hour, were continuously monitored and plotted in figure 1. It is worthwhile to note that the initial decrease of conversion is reversed after a few hours; thus the conversion of the olefins itself is likely to be a stimulating activation for the catalyst. In other words, catalyst and reaction effect each other by a mutual and beneficial activating action.

EXAMPLE 37

Example 36 was repeated by using a mixture 65% b.w. silicate-1 + 35% b.w. Al_2O_3 and by raising the space velocity to 67 h^{-1} . In this case, the conversion initially decreased and then the phenomenon was reversed.

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TABLE 1

EXAMPLE	1	2	3	4	5	6
Operative conditions:						
Catalyst	Boralite	see Ex. 1	Silicalite-1	See Ex. 3	Chromosilicate	See Ex. 5
SiO ₂ /Al ₂ O ₃ (moles)	(b)	"	∞	"	(d)	"
Amount of catalyst	3 grams (c)	"	see Ex. 1	"	see Ex. 1	"
Temp. (°C)	500	"	"	"	"	"
Space velocity (a)	6	"	"	"	"	"
Data survey after:	1 h	5 h	1 h	7 h	1 h	6 h
Results (2b.w.)						
Conversion	45.95	28.01	61.99	77.06	29.75	12.25
Selectivity to iso-C ₄	29.79	40.40	15.26	13.47	34.07	40.46
Selectivity to iso-C ₃	28.36	24.80	44.79	66.39	29.73	23.85
Yield (iso C ₄ + C ₃)	25.72	18.94	49.23	45.13	19.88	7.85
Sel. to compounds C ₄ -C ₅	38.76	31.73	28.15	31.75	32.29	30.82
Sel. to saturated compounds ≤ C ₄	2.51	2.55	6.85	5.44	3.46	4.77
Selectivity to C ₂	0.57	0.43	4.95	2.95	0.45	0.31

(a) WHSV (Weight Hourly Space Velocity) namely Kg/h of olefine per Kg. of pure catalyst (binder excluded);

(b) SiO₂:B₂O₃ = 7 (by moles);

(c) + 0.9 g of SiO₂ as binder;

(d) SiO₂:Cr₂O₃ = 36 (by moles)

TABLE 2

EXAMPLE	8	9	10	11	12	13	14	15
Operative Conditions:								
Catalyst	Silicalite-1	See Ex. 1	Boralite (d)	See Ex. 10	Chromosilicate (K)	See Ex. 12	Cr-Silicalite	See Ex. 14
SiO ₂ /Al ₂ O ₃ (moles)	∞	"	"	"	See Ex. 8	"	∞	"
Amount of catalyst	(c)	"	See Ex. 8	"	"	"	See Ex. 8	"
Temp. (°C)	550	"	"	"	"	"	"	"
Space velocity (a)	7.14	"	"	"	"	"	"	"
Data survey after:	1 h	6 h	1 h	6 h	1 h	6 h	1 h	6 h
Results (% b.w.)								
Conversion	64.26	57.00	42.81	19.52	16.80	10.04	69.22	60.92
Selectivity to C ₂ ⁻	19.44	19.80	23.43	27.97	21.25	17.92	20.51	19.63
Selectivity to C ₃ ⁻	33.19	39.42	39.27	38.93	32.50	29.62	36.25	39.71
Yield (C ₂ ⁻ + C ₃ ⁻)	52.63	59.22	62.70	66.90	53.75	47.54	56.76	59.34
Sel. to saturated compounds < C ₄	0.96	0.70	0.46	0.31	0.65	0.79	0.91	0.75
Sel. to C ₄ ⁺	29.27	29.42	27.61	24.70	26.90	30.38	29.31	28.37
Sel. to C ₄ ⁺	1.36	0.84	0.66	0.25	0.29	0.29	1.21	1.88
Sel. to BTX (b)	6.65	5.96	3.59	1.79	7.97	4.58	7.79	6.06
Sel. to others	3.06	3.82	4.95	5.79	10.83	7.27	6.02	3.69

(a) See ex. 1; (b) BTX = benzene + toluene + xylenes; (c) 0.5 g of pure zeolite + 0.21 g of Al₂O₃ (binder); (d) SiO₂/B₂O₃ = 7 (by moles);

(K) SiO₂ : Cr₂O₃ = 38 (by moles).

(*) Silicalite-1 containing 0.1% b.w. Cr, added by means of impregnation.

TABLE 3

EXAMPLE	16	17	18	19	20	21	22	23
Operative Conditions:	(*)	(*)	(**)	(**)	(**)	(***)	(***)	(***)
Catalyst	Silicalite-1	See Ex. 16	See Ex. 16	See Ex. 18	See Ex. 18	See Ex. 16	See Ex. 21	See Ex. 21
SiO ₂ /Al ₂ O ₃ (moles)	∞	"	"	"	"	"	"	"
Amount of catalyst	See Ex. 8	"	"	"	"	"	"	"
Temp. (°C)	550	"	"	"	"	"	"	600
Space velocity (s)	7.14	"	"	"	"	"	"	7.14
Data survey after:	1 h	6 h	1 h	6 h	7 h	1 h	6 h	7 h
Results (% b.w.)								
Conversion	98.01	96.83	97.79	97.36	89.9	100	97.47	95.50
Selectivity to C ₂ ⁻	10.38	9.12	10.55	9.29	7.48	9.16	1.60	16.64
Selectivity to C ₃ ⁻	51.94	52.90	53.03	53.29	55.35	31.64	33.08	27.55
Yield (C ₂ ⁻ + C ₃ ⁻)	60.28	60.05	52.17	61.41	56.48	40.80	37.75	42.21
Sel. to saturated compounds < C ₄	0.78	0.65	0.60	0.41	0.20	0.62	0.44	0.34
Sel. to C ₄ ⁺	17.55	14.80	16.78	13.65	11.92	26.96	28.21	24.37
Sel. to C ₄ ⁺	0.41	0.23	0.30	0.13	0.06	1.06	0.65	0.30
Sel. to BTX (a)	14.64	17.55	13.82	16.54	20.67	24.14	27.08	21.97
Sel. to C ₅	2.81	2.22	1.89	2.76	0.99	2.66	2.05	5.64
Sel. to others	2.20	2.51	2.97	3.40	3.26	3.74	2.77	3.14

(a) See ex. 1 (b) BTX = benzene + toluene + xylenes ; (*) Feed = n-hexene-1; (**) Feed = 2-methyl-pentene-1; (***) Feed = n-octene-1

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TABLE 4

EXAMPLE	24	25	26	27	28	29
<u>Operative conditions:</u>						
Catalyst	Silicalite-1	See Ex. 24	Boralite	See Ex. 26	Chromosilicate	See Ex. 28
$\text{SiO}_2/\text{Al}_2\text{O}_3$ (moles)	∞	"	(a)	"	(b)	"
Amount of catalyst	2.12 g	"	See Ex. 24	"	See Ex. 24	"
Temp. ($^{\circ}\text{C}$)	550	"	550	"	550	"
Space velocity (h^{-1})	7.5	"	7.5	"	7.5	"
Data survey after:	1 h	6 h	1 h	6 h	1 h	6 h
<u>Results (% b.w.)</u>						
Conversion	81.8	77.3	77.4	52.4	44.5	18.1
Selectivity to C_3	44.8	40.8	39.1	24.8	16.5	12.4
Yield (C_3)	36.5	31.5	30.3	15.8	7.4	2.2
Sel. to C_2	4.2	2.6	3.4	1.2	0.8	0.5
Sel. to saturated compounds $< \text{C}_4$	9.5	5.6	5.7	2.9	5.1	3.9
Sel. to compounds $> \text{C}_5$	6.4	7.0	7.8	5.4	13.7	8.1
Sel. to linear butenes	35.1	44.0	44.0	65.7	63.9	75.1

(a) $\text{SiO}_2 : \text{B}_2\text{O}_3 = 7$ (by moles); (b) $\text{SiO}_2 : \text{Cr}_2\text{O}_3 = 38$ (by moles).

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TABLE 5

EXAMPLE	32	33	34	35
	(*)	(*)	(**)	(**)
Operative conditions:	Silicalite-1	See Ex. 32	See Ex. 32	See Ex. 34
Catalyst	∞	"	"	"
SiO ₂ /Al ₂ O ₃ (moles)	3 g	"	"	"
Amount of catalyst	570	"	580	"
Temp., (°C)	60	"	60	"
Space velocity (h ⁻¹)	1 h	8 h	1 h	8 h
Data survey after:				
Results (% b.w.)				
Conversion	82.24	73.80	79.31	87.32
Selectivity to C ₃	31.98	39.89	39.66	26.00
Yield (C ₃)	26.30	29.44	31.45	17.50
Sel. to C ₂	3.89	3.19	4.21	3.20
Sel. to saturated compounds	13.35	10.62	14.14	8.40
<C _L				
Sel. to compounds >C ₅	36.02	25.72	27.80	45.60
Sel. to isobutene	14.76	20.58	14.19	16.50

(*) Pressure = 6 absolute atmospheres;

(**) Pressure = 9 absolute atmospheres.

CLAIMS

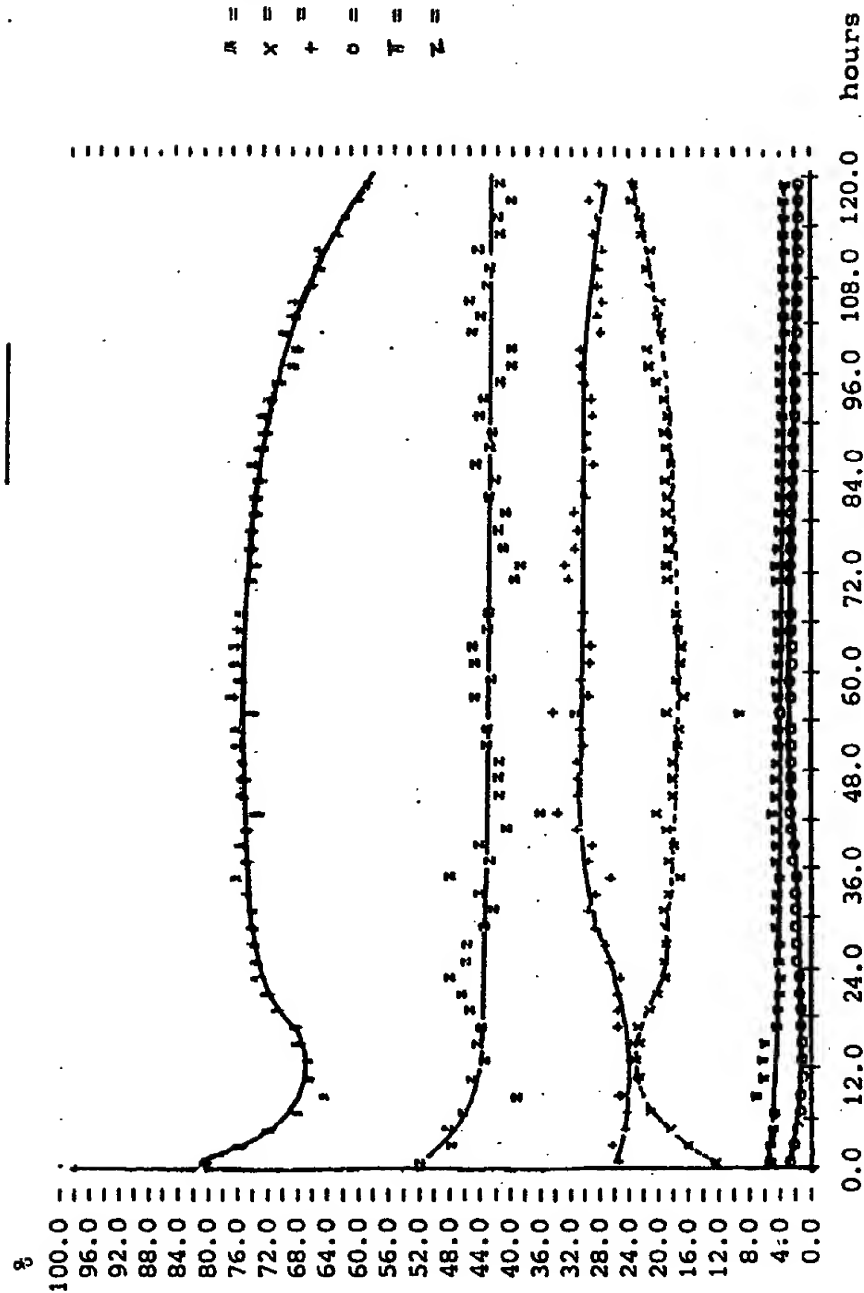
- 5 1. A process for the conversion of olefins having from 4 to 12 carbon atoms to propylene, characterized in that said olefins are brought into contact with a zeolitic compound (optionally in admixture with a binder), selected from silicalites, borasilites, chromosilicates and those
10 zeolites ZSM5 and ZSM11 in which $\text{SiO}_2/\text{Al}_2\text{O}_3$ (by moles) is ≥ 350 , at a space velocity of from 5 to 200 kg/h of olefins per kg of pure zeolitic compound (binder excluded) and at a temperature of from 400 to 600°C, said zeolitic compound being used as such or in a modified
15 form.
2. The process of claim 1, wherein said compound is silicalite-1, the pressure is substantially atmospheric and the space velocity is from 5 to 50 h^{-1} .
- 20 3. The process of claim 1, wherein said compound is silicalite-1, the pressure is from 1.5 to 7.5 absolute atmospheres and the space velocity is from 50 to 200 h^{-1} .
- 25 4. The process of any of claims 1-3, wherein said compound is silicalite-1 which is activated in the conversion reaction of ~~said olefins to propylene, under the~~ operative conditions of claim 1, the initial (activating) run being carried out at least 8 and preferably 12 hours.
- 30 5. The process of any of claims 1-4, wherein said olefins are selected from isobutene and linear butenes and wherein said binder is SiO_2 .
- 35 6. The process of any of claims 1-4, wherein said olefins contain from 5 to 8 carbon atoms and wherein said binder is Al_2O_3 .

7. The process of any of claims 1-6, wherein said compound is silicalite-1 in a non-modified form.
- 5 8. The process of any of claims 1-6, wherein said compound is silicalite-1 in a modified form, the modifying element being Cr, Mg, Ca, Sr or Ba.
- 10 9. A process for the conversion of olefins having 4 carbon atoms into propylene, characterized in that said olefins are brought into contact with a zeolitic compound (optionally in admixture with a binder) selected from silicalites and those zeolites ZSM5 and ZSM11 in which $\text{SiO}_2/\text{Al}_2\text{O}_3$ (by moles) is ≥ 350 , at a pressure of from 1.5 to 7.5 atmospheres, a space velocity of from 50 to 200 kg/h of olefins per kg of pure zeolitic compound (binder excluded) and a temperature of from 400 to 600°C, said compound being used as such or in a form modified by an element selected from Cr, Mg, Ca, Sr and Ba.
- 15 20 10. A process for the conversion of olefins having from 5 to 8 carbon atoms into propylene, characterized in that said olefins are brought into contact with a zeolitic compound (optionally in admixture with a binder) selected from silicalites and those zeolites ZSM5 and ZSM11 in which $\text{SiO}_2/\text{Al}_2\text{O}_3$ (by moles) is ≥ 350 , at a pressure of from 1.5 to 7.5 atmospheres, a space velocity of from 50 to 200 kg/h of olefins per kg of pure zeolitic compound (binder excluded) and a temperature of from 400 to 600°C, said compound being used as such or in a form modified by an element selected from Cr, Mg, Ca, Sr and Ba.
- 25 30 11. The process of claims 9 or 10 wherein said compound is silicalite-1 which is activated in the conversion reaction of said olefins to propylene, under said operative conditions, the initial (activating) run being carried out at least 8 and preferably 12 hours.
- 35

12. A process for the conversion into propylene of mixtures
of olefins having 4 C atoms, said mixtures containing
also C_4^+ paraffins that cannot be easily separated from
olefins, which process comprises the following steps:

- a) preliminary oligomerization of a $(C_4^- + C_4^+)$ mixture
at a temperature of from 320 to 380°C using a catalyst
bed of zeolitic nature to obtain a mixture of olefins
having from 5 to 8 C atoms, with the C_4^+ paraffins
remaining unconverted;
- b) cooling and condensation of the oligomerization
effluent in order to separate the C_4^+ paraffins as a
gaseous phase, and conversion of the remaining
 $(C_5^- - C_8^-)$ mixture to propylene under the typical
"post-pyrolysis" reaction conditions of claim 1;
- c) cooling the effluent from the reactor for the con-
version to propylene and compression of said effluent,
preferably at 13-16 absolute kg/cm^2 , whereby the
hydrocarbons having 4 or more C atoms are condensed
and the hydrocarbons having less than 4 C atoms
are separated as a gaseous phase.

FIGURE 1



A = Conversion
 x = Selectivity to iso-C₄
 + = Sel. to C₃
 o = Sel. to C₂
 r = Sel. to saturated gases
 z = Sel. to liquids \geq C₅

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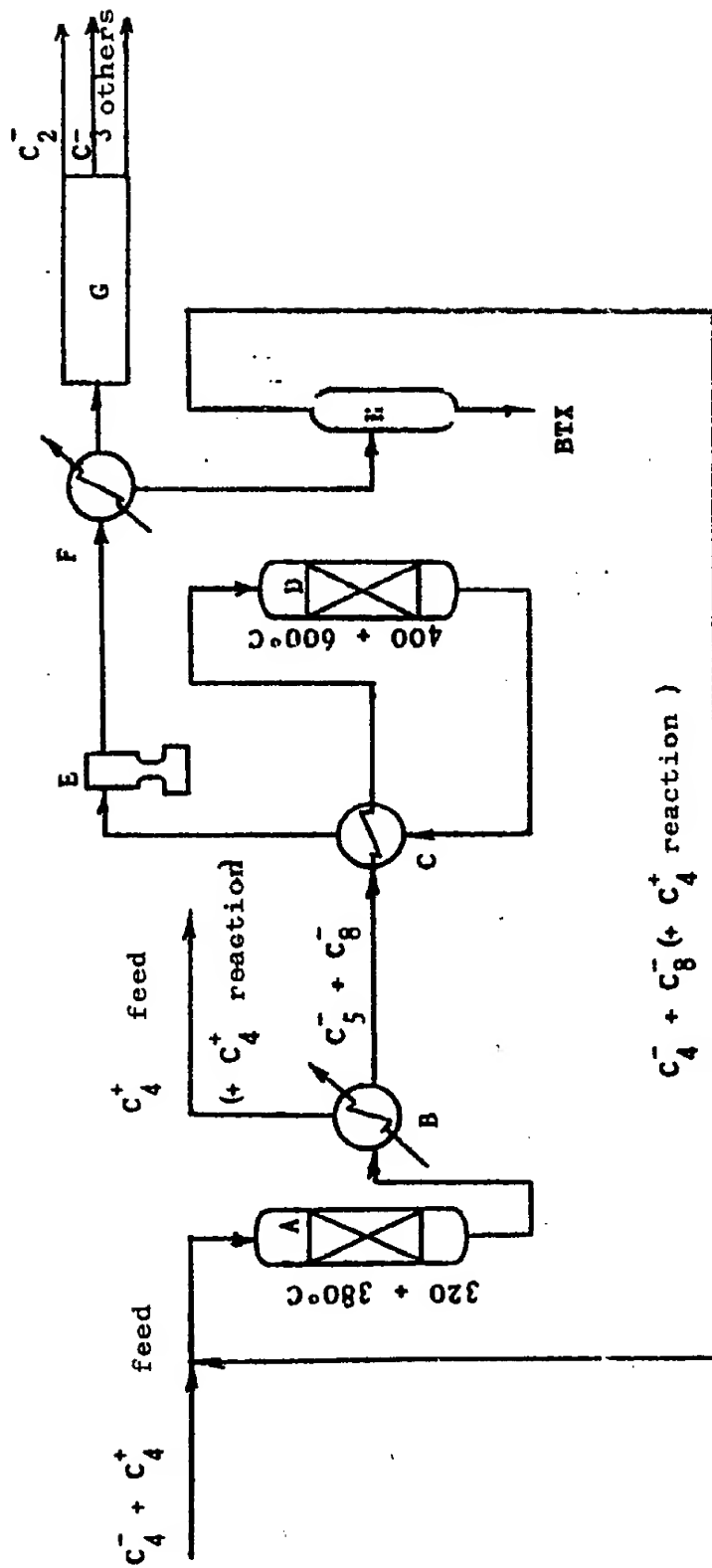


FIGURE 2



European Patent
Office

EUROPEAN SEARCH REPORT

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Application number

EP 83 11 1241

DOCUMENTS CONSIDERED TO BE RELEVANT			
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1	A,D US-A-4 150 062 (GARWOOD et al.) * Examples 5-6; claims *	12	
			TECHNICAL FIELDS SEARCHED (Int. Cl. 7)
			C 07 C 11/00 C 07 C 4/00
The present search report has been drawn up for all claims			
Place of search THE HAGUE		Date of completion of the search 30-01-1984	Examiner VAN GEYT J.J.A.
CATEGORY OF CITED DOCUMENTS			
X : particularly relevant if taken alone Y : particularly relevant if combined with another document of the same category A : technological background O : non-written disclosure P : intermediate document			
T : theory or principle underlying the invention E : earlier patent document, but published on, or after the filing date D : document cited in the application L : document cited for other reasons & : member of the same patent family, corresponding document			